ANSWERS

Chapter 8

8.1 1.8

8.2 (a) From the given graph for a stress of 150×10^6 N m^{-2} the strain is 0.002

(b) Approximate yield strength of the material is 3×10^8 N m^{-2}

8.3 (a) Material A

(b) Strength of a material is determined by the amount of stress required to cause fracture: material A is stronger than material B.

8.4 (a) False (b) True

8.5 1.5×10^{-4} m (steel); 1.3×10^{-4} m (brass)

8.6 Deflection = 4×10^{-6} m

8.7 2.8×10^{-6}

8.8 0.127

8.9 7.07×10^4 N

8.10 $D_{copper}/D_{iron} = 1.25$

8.11 1.539×10^{-4} m

8.12 2.026×10^9 Pa

8.13 1.034×10^3 kg/m^3

8.14 0.0027

8.15 0.058 cm^3

8.16 2.2×10^6 N/m^2

Chapter 9

9.3 (a) decreases (b) η of gases increases, η of liquid decreases with temperature (c) shear strain, rate of shear strain (d) conservation of mass, Bernoulli's equation (e) greater.

9.5 6.2×10^6 Pa

9.6 10.5 m

19.7 Pressure at that depth in the sea is about 3×10^7 Pa. The structure is suitable since it can withstand far greater pressure or stress.

9.8 6.92×10^5 Pa

9.9 0.800

9.10 Mercury will rise in the arm containing spirit; the difference in levels of mercury will be 0.221 cm.

9.11 No, Bernoulli's principle applies to streamline flow only.

9.12 No, unless the atmospheric pressures at the two points where Bernoulli's equation is applied are significantly different.

9.13 9.8×10^2 Pa (The Reynolds number is about 0.3 so the flow is laminar).

9.14 1.5×10^3 N

9.15 Fig (a) is incorrect [Reason: at a constriction (i.e. where the area of cross-section of the tube is smaller), flow speed is larger due to mass conservation. Consequently pressure there is smaller according to Bernoulli's equation. We assume the fluid to be incompressible].

9.16 0.64 m s^{-1}

9.17 2.5×10^{-2} N m^{-1}

9.18 4.5×10^{-2} N for (b) and (c), the same as in (a).

9.19 Excess pressure = 310 Pa, total pressure = 1.0131×10^5 Pa. However, since data are correct to three significant figures, we should write total pressure inside the drop as 1.01×10^5 Pa.

9.20 Excess pressure inside the soap bubble = 20.0 Pa; excess pressure inside the air bubble in soap solution = 10.0 Pa. Outside pressure for air bubble = $1.01 \times 10^5 + 0.4 \times 10^3 \times 9.8 \times 1.2 = 1.06 \times 10^5$ Pa. The excess pressure is so small that up to three significant figures, total pressure inside the air bubble is 1.06×10^5 Pa.

Chapter 10

10.1 Neon: – 248.58 °C = – 415.44 °F;

CO_2: – 56.60 °C = – 69.88 °F

(use $t_F = \frac{9}{5}t_c + 32$)

10.2 $T_A = (4/7)\ T_B$

10.3 384.8 K

10.4 (a) Triple-point has a *unique* temperature; fusion point and boiling point temperatures depend on pressure; (b) The other fixed point is the absolute zero itself; (c) Triple-point is 0.01°C, not 0 °C; (d) 491.69.

10.5 (a) T_A = 392.69 K, T_B = 391.98 K; (b) The discrepancy arises because the gases are not perfectly ideal. To reduce the discrepancy, readings should be taken for lower and lower pressures and the plot between temperature measured versus absolute pressure of the gas at triple point should be extrapolated to obtain temperature in the limit pressure tends to zero, when the gases approach ideal gas behaviour.

10.6 Actual length of the rod at 45.0 °C = (63.0 + 0.0136) cm = 63.0136 cm. (However, we should say that change in length up to three significant figures is 0.0136 cm, but the total length is 63.0 cm, up to three significant places. Length of the same rod at 27.0 °C = 63.0 cm.

10.7 When the shaft is cooled to temperature – 69°C the wheel can slip on the shaft.

10.8 The diameter increases by an amount = 1.44×10^{-2} cm.

10.9 3.8×10^2 N

10.10 Since the ends of the combined rod are not clamped, each rod expands freely.

Δl_{brass} = 0.21 cm, Δl_{steel} = 0.126 cm = 0.13 cm

Total change in length = 0.34 cm. No 'thermal stress' is developed at the junction since the rods freely expand.

10.11 0.0147 = 1.5×10^{-2}

10.12 103 °C

10.13 1.5 kg

10.14 0.43 J g^{-1} K^{-1} ; smaller

10.15 The gases are diatomic, and have other degrees of freedom (i.e. have other modes of motion) possible besides the translational degrees of freedom. To raise the temperature of the gas by a certain amount, heat is to be supplied to increase the average energy of all the modes. Consequently, molar specific heat of diatomic gases is more than that of monatomic gases. It can be shown that if only rotational modes of motion are considered, the molar specific heat of diatomic gases is nearly (5/2) R which agrees with the observations for all the gases listed in the table, except chlorine. The higher value of molar specific heat of chlorine indicates that besides rotational modes, vibrational modes are also present in chlorine at room temperature.

10.16 4.3 g/min

10.17 3.7 kg

10.18 238 °C

10.20 9 min

Chapter 11

11.1 16 g per min

11.2 934 J

11.4 2.64

11.5 16.9 J

11.6 (a) 0.5 atm (b) zero (c) zero (assuming the gas to be ideal) (d) No, since the process (called free expansion) is rapid and cannot be controlled. The intermediate states are non-equilibrium states and do not satisfy the gas equation. In due course, the gas does return to an equilibrium state.

11.7 25 W

11.8 450 J

Chapter 12

12.1 4×10^{-4}

12.3 (a) The dotted plot corresponds to 'ideal' gas behaviour; (b) $T_1 > T_2$; (c) 0.26 J K^{-1}; (d) No, 6.3×10^{-5} kg of H_2 would yield the same value

12.4 0.14 kg

12.5 5.3×10^{-6} m^3

12.6 6.10×10^{26}

12.7 (a) 6.2×10^{-21} J (b) 1.24×10^{-19} J (c) 2.1×10^{-16} J

12.8 Yes, according to Avogadro's law. No, v_{rms} is largest for the lightest of the three gases; neon.

12.9 2.52×10^3 K

12.10 Use the formula for mean free path :

$$\bar{l} = \frac{1}{\sqrt{2}\pi n d^2}$$

where d is the diameter of a molecule. For the given pressure and temperature $N/V = 5.10 \times 10^{25}$ m^{-3} and $= 1.0 \times 10^{-7}$ m. $v_{rms} = 5.1 \times 10^2$ m s^{-1}.

collisional frequency $= \frac{v_{rms}}{\bar{l}} = 5.1 \times 10^9 \, s^{-1}$. Time taken for the collision $= d / v_{rms} = 4 \times 10^{-13}$ s.

Time taken between successive collisions $= l / v_{rms} = 2 \times 10^{-10}$ s. Thus the time taken between successive collisions is 500 times the time taken for a collision. Thus a molecule in a gas moves essentially free for most of the time.

Chapter 13

13.1 (b), (c)

13.2 (b) and (c): SHM; (a) and (d) represent periodic but not SHM [A polyatomic molecule has a number of natural frequencies; so in general, its vibration is a superposition of SHM's of a number of different frequencies. This superposition is periodic but not SHM].

13.3 (b) and (d) are periodic, each with a period of 2 s; (a) and (c) are not periodic. [Note in (c), repetition of merely one position is not enough for motion to be periodic; the entire motion during one period must be repeated successively].

13.4 (a) Simple harmonic, $T = (2\pi/\omega)$; (b) periodic, $T = (2\pi/\omega)$ but not simple harmonic; (c) simple harmonic, $T = (\pi/\omega)$; (d) periodic, $T = (2\pi/\omega)$ but not simple harmonic; (e) non-periodic; (f) non-periodic (physically not acceptable as the function $\to \infty$ as $t \to \infty$.

13.5 (a) 0, +, + ; (b) 0, –, – ; (c) –, 0,0 ; (d) –, –, – ; (e) +, +, + ; (f) –, –, –.

13.6 (c) represents a simple harmonic motion.

13.7 A = $\sqrt{2}$ cm, $\phi = 7\pi/4$; B = $\sqrt{2}$ cm, $a = \pi/4$.

13.8 219 N

13.9 Frequency 3.2 s^{-1}; maximum acceleration of the mass 8.0 m s^{-2}; maximum speed of the mass 0.4 m s^{-1}.

13.10 (a) $x = 2 \sin 20t$
(b) $x = 2 \cos 20t$
(c) $x = -2 \cos 20t$

where x is in cm. These functions differ neither in amplitude nor frequency. They differ in initial phase.

13.11 (a) $x = -3 \sin \pi t$ where x is in cm.

(b) $x = -2 \cos \frac{\pi}{2} t$ where x is in cm.

13.13 (a) F/k for both (a) and (b).

(b) $T = 2\pi\sqrt{\frac{m}{k}}$ for (a) and $2\pi\sqrt{\frac{m}{2k}}$ for (b)

13.14 100 m/min

13.15 8.4 s

13.16 $T = 2\pi\sqrt{\dfrac{l}{\sqrt{g^2 + v^4/R^2}}}$. Hint: Effective acceleration due to gravity will get reduced due to radial acceleration v^2/R acting in the horizontal plane.

13.17 In equilibrium, weight of the cork equals the up thrust. When the cork is depressed by an amount x, the net upward force is $Ax\rho_l g$. Thus the force constant $k = A\rho_l g$.

Using $m = Ah\rho$, and $T = 2\pi\sqrt{\dfrac{m}{k}}$ one gets the given expression.

13.18 When both the ends are open to the atmosphere, and the difference in levels of the liquid in the two arms is h, the net force on the liquid column is $Ah\rho g$ where A is the area of cross-section of the tube and ρ is the density of the liquid. Since restoring force is proportional to h, motion is simple harmonic.

Chapter 14

14.1 0.5 s

14.2 8.7 s

14.3 2.06×10^4 N

14.4 Assume ideal gas law: $P = \dfrac{\rho RT}{M}$, where ρ is the density, M is the molecular mass, and T is the temperature of the gas. This gives $v = \sqrt{\dfrac{\gamma RT}{M}}$. This shows that v is:

(a) Independent of pressure.

(b) Increases as $\sqrt{T}$.

(c) The molecular mass of water (18) is less than that of N_2 (28) and O_2 (32).

Therefore as humidity increases, the effective molecular mass of air decreases and hence v increases.

14.5 The converse is not true. An obvious requirement for an acceptable function for a travelling wave is that it should be finite everywhere and at all times. Only function (c) satisfies this condition, the remaining functions cannot possibly represent a travelling wave.

14.6 (a) 3.4×10^{-4} m (b) 1.49×10^{-3} m

14.7 4.1×10^{-4} m

14.8 (a) A travelling wave. It travels from right to left with a speed of 20 ms^{-1}.

(b) 3.0 cm, 5.7 Hz

(c) $\pi/4$

(d) 3.5 m

14.9 All the graphs are sinusoidal. They have same amplitude and frequency, but different initial phases.

14.10 (a) 6.4 π rad

(b) 0.8 π rad

(c) π rad

(d) $(\pi/2)$ rad

14.11 (a) Stationary wave

(b) l = 3 m, n = 60 Hz, and v = 180 m s^{-1} for each wave

(c) 648 N

14.12 (a) All the points except the nodes on the string have the same frequency and phase, but not the same amplitude.

(b) 0.042 m

14.13 (a) Stationary wave.

(b) Unacceptable function for any wave.

(c) Travelling harmonic wave.

(d) Superposition of two stationary waves.

14.14 (a) 79 m s^{-1}

(b) 248 N

14.15 347 m s^{-1}

Hint : $v_n = \frac{(2n-1)v}{4l}$; n = 1,2,3,….for a pipe with one end closed

14.16 5.06 km s^{-1}

14.17 First harmonic (fundamental); No.

14.18 318 Hz

BIBLIOGRAPHY

TEXTBOOKS

For additional reading on the topics covered in this book, you may like to consult one or more of the following books. Some of these books however are more advanced and contain many more topics than this book.

1. **Ordinary Level Physics**, A.F. Abbott, Arnold-Heinemann (1984).
2. **Advanced Level Physics**, M. Nelkon and P. Parker, 6th Edition Arnold-Heinemann (1987).
3. **Advanced Physics**, Tom Duncan, John Murray (2000).
4. **Fundamentals of Physics**, David Halliday, Robert Resnick and Jearl Walker, 7th Edition John Wily (2004).
5. **University Physics**, H.D. Young, M.W. Zemansky and F.W. Sears, Narosa Pub. House (1982).
6. **Problems in Elementary Physics**, B. Bukhovtsa, V. Krivchenkov, G. Myakishev and V. Shalnov, MIR Publishers, (1971).
7. **Lectures on Physics** (3 volumes), R.P. Feynman, Addision – Wesley (1965).
8. **Berkeley Physics Course** (5 volumes) McGraw Hill (1965).
   a. Vol. 1 – Mechanics: (Kittel, Knight and Ruderman)
   b. Vol. 2 – Electricity and Magnetism (E.M. Purcell)
   c. Vol. 3 – Waves and Oscillations (Frank S. Craw-ford)
   d. Vol. 4 – Quantum Physics (Wichmann)
   e. Vol. 5 – Statistical Physics (F. Reif)
9. **Fundamental University Physics**, M. Alonso and E. J. Finn, Addison – Wesley (1967).
10. **College Physics**, R.L. Weber, K.V. Manning, M.W. White and G.A. Weygand, Tata McGraw Hill (1977).
11. **Physics**: **Foundations and Frontiers**, G. Gamow and J.M. Cleveland, Tata McGraw Hill (1978).
12. **Physics for the Inquiring Mind**, E.M. Rogers, Princeton University Press (1960)
13. **PSSC Physics Course**, DC Heath and Co. (1965) Indian Edition, NCERT (1967)
14. **Physics Advanced Level**, Jim Breithampt, Stanley Thornes Publishers (2000).
15. **Physics,** Patrick Fullick, Heinemann (2000).

16. **Conceptual Physics,** Paul G. Hewitt, Addision-Wesley (1998).
17. **College Physics,** Raymond A. Serway and Jerry S. Faughn, Harcourt Brace and Co. (1999).
18. **University Physics,** Harris Benson, John Wiley (1996).
19. **University Physics,** William P. Crummet and Arthur B. Western, Wm.C. Brown (1994).
20. **General Physics,** Morton M. Sternheim and Joseph W. Kane, John Wiley (1988).
21. **Physics,** Hans C. Ohanian, W.W. Norton (1989).
22. **Advanced Physics,** Keith Gibbs, Cambridge University Press(1996).
23. **Understanding Basic Mechanics,** F. Reif, John Wiley (1995).
24. **College Physics,** Jerry D. Wilson and Anthony J. Buffa, Prentice-Hall (1997).
25. **Senior Physics, Part – I,** I.K. Kikoin and A.K. Kikoin, Mir Publishers (1987).
26. **Senior Physics, Part – II,** B. Bekhovtsev, Mir Publishers (1988).
27. **Understanding Physics,** K. Cummings, Patrick J. Cooney, Priscilla W. Laws and Edward F. Redish, John Wiley (2005)
28. **Essentials of Physics,** John D. Cutnell and Kenneth W. Johnson, John Wiley (2005)

GENERAL BOOKS

For instructive and entertaining general reading on science, you may like to read some of the following books. Remember however, that many of these books are written at a level far beyond the level of the present book.

1. **Mr. Tompkins** in paperback, G. Gamow, Cambridge University Press (1967).
2. **The Universe and Dr. Einstein**, C. Barnett, Time Inc. New York (1962).
3. **Thirty years that Shook Physics**, G. Gamow, Double Day, New York (1966).
4. **Surely You're Joking, Mr. Feynman**, R.P. Feynman, Bantam books (1986).
5. **One, Two, Three... Infinity**, G. Gamow, Viking Inc. (1961).
6. **The Meaning of Relativity**, A. Einstein, (Indian Edition) Oxford and IBH Pub. Co (1965).
7. **Atomic Theory and the Description of Nature**, Niels Bohr, Cambridge (1934).
8. **The Physical Principles of Quantum Theory**, W. Heisenberg, University of Chicago Press (1930).
9. **The Physics- Astronomy Frontier**, F. Hoyle and J.V. Narlikar, W.H. Freeman (1980).
10. **The Flying Circus of Physics with Answer**, J. Walker, John Wiley and Sons (1977).
11. **Physics for Everyone** (series), L.D. Landau and A.I. Kitaigorodski, MIR Publisher (1978).
    Book 1: Physical Bodies
    Book 2: Molecules
    Book 3: Electrons
    Book 4: Photons and Nuclei.
12. **Physics can be Fun**, Y. Perelman, MIR Publishers (1986).
13. **Power of Ten**, Philip Morrison and Eames, W.H. Freeman (1985).
14. **Physics in your Kitchen Lab**., I.K. Kikoin, MIR Publishers (1985).
15. **How Things Work : The Physics of Everyday Life**, Louis A. Bloomfield, John Wiley (2005)
16. **Physics Matters : An Introduction to Conceptual Physics,** James Trefil and Robert M. Hazen, John Wiley (2004).

INDEX

Notes